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⑭ **Polycondensed polysilazane and method of producing same.**

⑮ A novel, polycondensed polysilazane is obtained by reacting a polysilazane with a compound selected from  
secondary amines and tri-substituted hydrazines so as to link the compound to the polysilazane.

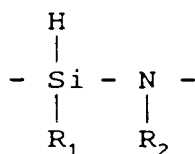
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This invention relates to a novel polysilazane and to a method of preparing same by polycondensation.

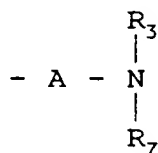
There are known polysilazanes which are proposed to be used as a precursor material for the production of silicon nitride-containing ceramics. Known polysilazanes are liquid or solid and the solid polysilazanes are insoluble in an organic solvent such as toluene. For example, A. Stock discloses an inorganic polysilazane of the formula  $(\text{SiH}_2\text{NH})_n$  prepared by reacting dichlorosilane with ammonia using benzene as a solvent (Ber. 54, 740 (1921)). This inorganic polysilazane is an oligomer ( $n = 7$  to 8) and is a viscous liquid at room temperature. D. Seyferth et al suggest an inorganic polysilazane obtained by reacting dichlorosilane with ammonia using dichloromethane as a solvent (US-A-4397828). This polysilazane is an oily liquid having a proton ratio Si-H/N-H of about 3.3 and becomes solidified when heated at about 200°C or allowed to stand for 3-5 days. The solidified polysilazane is insoluble in an organic solvent such as o-xylene. Japanese Published Unexamined Patent Application (Tokkyo Kokai) No. 60-145,903 discloses an inorganic polysilazane having a molecular weight of 690-2,000 and obtained by reacting a dihalosilane adduct such as a dichlorosilanepyridine adduct with ammonia. The resultant polysilazane upon removal of the solvent therefrom is a viscous liquid or a resinous solid. This solid, however, is insoluble in an organic solvent such as o-xylene.

In utilizing an inorganic polysilazane as raw materials for ceramic fibers, binders, coating agents or the like, it is highly desirable that the polysilazane be soluble in an organic solvent, have a high molecular weight and exhibit good spinnability, solidifiability and moldability. In these respects, the above-described known inorganic polysilazanes are not fully satisfactory.

The present invention has been made with the foregoing problems of the conventional inorganic polysilazane in view and provides a polycondensed polysilazane having a number average molecular weight of 200-500,000 and containing recurring units of the formula:

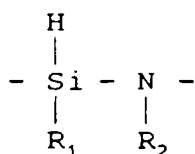


wherein  $\text{R}_1$  and  $\text{R}_2$ , independently from each other, stand for hydrogen, alkyl, alkenyl, cycloalkyl, alkylamino, aryl or alkylsilyl, and (b) one or more groups replacing hydrogen on and pendant from the silicon atoms of said repeating units, of the formula:

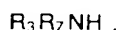


wherein  $\text{R}_1$  and  $\text{R}_2$  are as defined above. A stands for a direct bond or  $-\text{N}(\text{R}_4)-$  where  $\text{R}_4$  stands for hydrogen, alkyl, cycloalkyl, aryl, aralkyl or a heterocyclic group, and  $\text{R}_3$  and  $\text{R}_7$ , independently from each other, stand for alkyl, alkenyl, cycloalkyl, aryl, aralkyl or a heterocyclic group.

In another aspect, the present invention provides a method of forming a novel polycondensed polysilazane as defined above, comprising reacting a starting polysilazane having a number average molecular weight of 100-50,000 and containing recurring units of the formula:



wherein  $R_1$  and  $R_2$  are as defined above, in an organic base-containing solvent, with a secondary amine of the formula:



or with a 1,1,2-trisubstituted hydrazone of the formula:

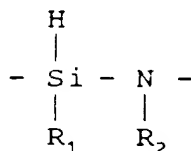


wherein  $R_3$ ,  $R_4$  and  $R_7$  are as defined above.

The groups represented by the symbols  $R_3$ ,  $R_4$  and  $R_7$ , other than  $R_4$  being hydrogen, may optionally be substituted by one or more substituents, which are inert to the protons on the silicon atoms of the polysilazane skeleton, such as alkyl, aryl, alkoxy and alkoxy carbonyl groups.

The present invention will now be described in greater detail below.

The raw material inorganic polysilazane to be used as the starting polysilazane in the method according to the present invention has as its main skeletal structure recurring units represented by the following formula:



wherein  $R_1$  and  $R_2$ , independently from each other, stand for hydrogen, alkyl such as methyl, ethyl, propyl, butyl, octyl or decyl; alkenyl such as vinyl, allyl, butenyl, octenyl or decenyl; cycloalkyl such as cyclohexyl or methylcyclohexyl; alkylamino such as methylamino or ethylamino; aryl such as phenyl, tolyl, xylyl or naphthyl; aralkyl such as benzyl; or alkylsilyl such as methylsilyl, ethylsilyl, propylsilyl, butylsilyl, octylsilyl or decylsilyl. These alkyl, alkenyl, cycloalkyl, aryl, alkyl amino, aralkyl and alkylsilyl groups may each optionally be substituted by one or more substituents, which are inert to the protons bonded to the silicon atoms of the polysilazane skeleton, such as an alkyl, an aryl, an alkoxy and an alkoxy carbonyl.

A raw material polysilazane having a number average molecular weight of 100-50,000, preferably 300-2,000, more preferably 600-1,400 is used. The polysilazane may have a cyclic or linear form. Polysilazanes having both cyclic and linear portions may also be suitably used. Examples of suitable raw material polysilazanes to be used for the purpose of the present invention are those obtained by the methods disclosed in:

Japanese Tokyo Kokai No. 60-145903;  
U.S. Patent No. 4,397,828 (Seyferth et al);  
Ber. 54, (1921), p 740 (Stock);  
Inorg. Chem. (1972), 11 (Scantlin et al);  
U.S. Patent No. 3,318,823 (Aylett);  
U.S. Patent No. 4,482,669 (Seyferth et al);  
Japanese Tokyo Kokai No. 61-89230.

The above polysilazane is reacted with a secondary amine of the formula:  $R_3R_7NH$ , or with a 1,1,2-trisubstituted hydrazine of the formula:  $R_3R_7N-NHR_4$ , wherein  $R_3$ ,  $R_4$  and  $R_7$  are as defined above.

Illustrative of suitable secondary amines are dimethylamine, diethylamine, methylethylamine, dipropylamine, diisopropylamine, methylpropylamine, dibutylamine, diamylamine, dihexylamine, diheptylamine, dioctylamine, diallylamine, allylcyclohexylamine, methylaniline, ethylaniline, dibenzylamine, diphenylamine, dicyclohexylamine and N-cyclohexylaniline.

Illustrative of suitable 1,1,2-tri-substituted hydrazines are trimethylhydrazine, triethylhydrazine, tripropylhydrazine, triphenylhydrazine, 1,2-dimethyl-1-phenylhydrazine, 1,1-dimethyl-2-phenylhydrazine and benzilidenemethylhydrazine.

In carrying out the polycondensation of the polysilazane with secondary amine or trisubstituted hydrazine, these materials are dissolved in an organic base-containing solvent and the solution is subjected to dehydrogenative polycondensation conditions. The organic base-containing solvent may be either a liquid organic base or a non-basic organic solvent having dissolved therein an organic base.

Any liquid organic base which does not react with the raw material polysilazane may be used as the organic base-containing solvent. Illustrative of suitable liquid organic bases are tertiary amines such as trimethylamine, dimethylamine, dimethylethylamine, diethylmethylamine, triethylamine, pyridine and a substituted pyridine, dimethylaniline and a substituted dimethylaniline, pyrazine and a substituted pyrazine, pyrimidine and a substituted pyrimidine, pyridazine and a substituted pyridazine, pyrrole, 3-pyrroline, pyrazole, 2-pyrazoline and mixtures thereof.

The organic base to be dissolved in the non-basic organic solvent may be tertiary amines such as described immediately above, secondary amines having bulky groups and phosphines. The non-basic organic solvent may be, for example, aliphatic hydrocarbons, alicyclic hydrocarbons, aromatic hydrocarbons, halogenated hydrocarbons, aliphatic ethers and alicyclic ethers. Illustrative of suitable non-basic organic solvents are methylene chloride, chloroform, carbon tetrachloride, bromoform, ethylene chloride, ethylidene chloride, trichloroethane, tetrachloroethane, ethyl ether, isopropyl ether, ethyl butyl ether, butyl ether, 1,2-dioxethane, dioxane, dimethyl dioxane, tetrahydrofuran, tetrahydropyran, pentane, hexane, isohexane, methylpentane, heptane, isohexane, octane, isooctane, cyclopentane, methylcyclopentane, cyclohexane, methylcyclohexane, benzene, toluene, xylene, ethylbenzene and mixtures thereof.

The amount of the organic base is at least 5 parts by weight, preferably at least 20% by weight per 100 parts by weight of the non-basic organic solvent.

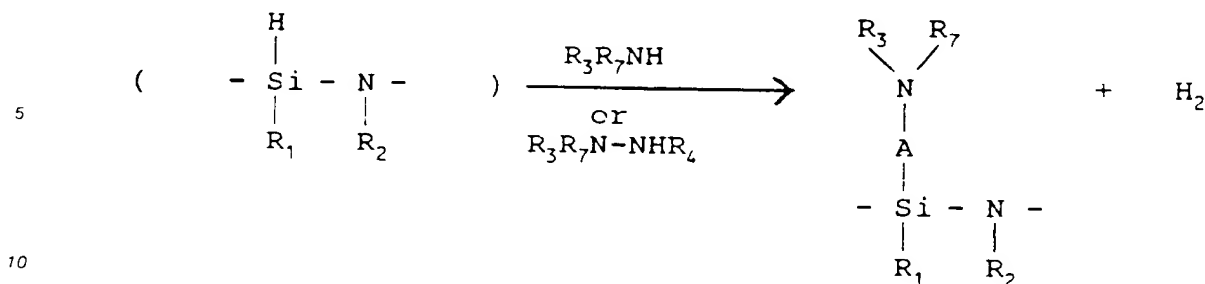
The solution of the raw material polysilazane in the organic base-containing solvent generally has a concentration of 0.1-50% by weight. Too high a concentration of the polysilazane in excess of 50% by weight causes a difficulty in controlling the reaction. On the other hand, too low a concentration below 0.1% by weight is undesirable because the reaction proceeds too slowly. The concentration of the raw material polysilazane is preferably 1-12% by weight.

The secondary amine or trisubstituted hydrazine may be used generally in an amount of 0.01-5000 mols, preferably 0.5-1000 mols per mole of the polysilazane.

The reaction is performed at a temperature of generally -78 to 300°C, preferably -48 to 200°C. A temperature below -78°C is insufficient to effect the polycondensation reaction while too high a temperature in excess of 300°C causes difficulties in homogeneously proceeding the reaction. Preferably the reaction is carried out in an atmosphere of dry nitrogen, dry argon or the like inert atmosphere. When the reaction is performed in a closed reactor such as autoclave, the reaction pressure increases as the reaction proceeds because of the in situ production of hydrogen gas. It is not necessary, however, to carry out the polycondensation in a pressurized condition. The reaction may be performed under ambient pressure. The reaction time varies with the kinds and concentrations of the raw material polysilazane and the organic base, the reaction temperature adopted and the intended properties of the polycondensed polysilazane product but generally in the range of about 0.5-20 hours. The optimum reaction conditions vary with the average molecular weight and molecular weight distribution of the raw material polysilazane. More severe conditions are generally adopted as the molecular weight of the raw material polysilazane becomes low.

The reaction mixture after the completion of the polycondensation reaction is generally a solution containing the polycondensed polysilazane and the organic base-containing solvent. It is desirable to reduce the concentration of the organic base in the reaction product, since otherwise a gelation of the reaction mixture will result when it is allowed to stand for a long time at room temperature. The reduction of the concentration of the organic base may be effected by removing at least a portion of the organic base by distillation and substituting therefor a suitable amount of a non-basic organic solvent such as a hydrocarbon, a halogenated hydrocarbon or an ether. The non-basic organic solvents exemplified previously are suitably used. Such a replacement operation may be repeated twice or more, if desired, to obtain a stable solution of the reformed polysilazane. In particular, it is preferred that the concentration of the organic base in the product polysilazane solution be 30% or less, more preferably 5% or less based on the total weight of the organic base and the non-basic organic solvent contained in the solution.

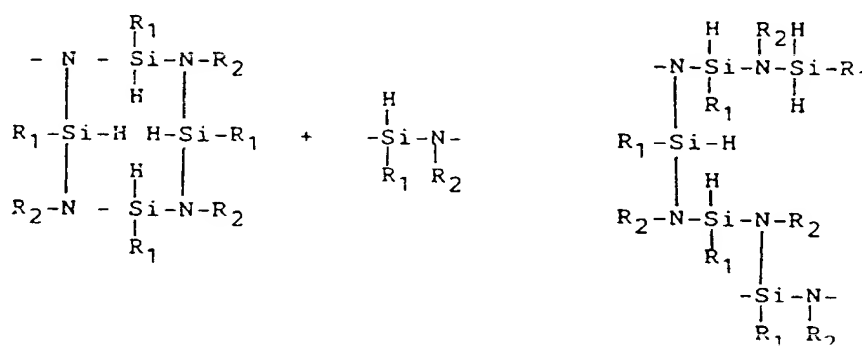
The reaction between the secondary amine or tri-substituted hydrazine and the starting polysilazane is a dehydrogenative condensation reaction which can be schematically shown as follows:



The polysilazane according to the present invention is a polymer obtained by the polycondensation reaction of the raw material polysilazane and the amine or hydrazine and which has newly introduced bonding with the group  $-\text{A}-\text{NR}_3\text{R}_7$ . Thus, the polycondensed polysilazane has a larger molecular weight than the raw material polysilazane. The polycondensed polysilazane has a molecular weight of 200-500,000, preferably 1,000-50,000, more preferably 1,500-10,000.

In addition, the polycondensed polysilazane has a higher content of nitrogen atoms than the raw material polysilazane. For instance, the atomic ratios N/Si of perhydropolysilazane, methylhydrosilazane, N-methylsilazane, N-(triethylsilyl)allylsilazane, N-(dimethylamino)cyclohexylsilazane and phenylpolysilazane are increased upon being converted into the reformed polysilazanes, typically from 0.60-0.75 to 0.80 or more, from 0.90-0.97 to 0.98 or more, from 0.67-1.50 to 1.6 or more, from 0.55-0.70 to 0.87 or more, from 1.1-2.0 to 2.2 or more, and from 0.85-0.96 to 0.98 or more, respectively. The N/Si ratio of the polycondensed polysilazane may be increased to any extent as long as the solubility of the polycondensed polysilazane in an organic solvent such as o-xylene is not adversely affected. Generally, the N/Si atomic ratio of the polycondensed polysilazane is not greater than 2.5, preferably not greater than 2.0.

The polycondensed polysilazane is also distinct from the raw material polysilazane in an increased number of branched side chains. For this reason, even when the polycondensation results in the formation of a polycondensed polysilazane which becomes solidified upon removal of the solvent, this solid is soluble in an organic solvent such as o-xylene, notwithstanding its high molecular weight. The branched chains are considered to result from the following reaction:



When the group  $\text{R}_1$  of the raw material polysilazane is hydrogen, the formation of the branched chains may be confirmed by comparing proton NMR spectra of the polycondensed polysilazane and the raw material polysilazane. Namely, the integration ( $\text{S}_{2\text{H}}$ ) of the peak at  $\delta 4.8$  ppm (attributed to the  $\text{SiH}_2$  group) and that ( $\text{S}_{3\text{H}}$ ) of the peak at  $\delta 4.4$  ppm (attributed to the  $\text{SiH}_3$  group) gives a  $\text{SiH}_2/\text{SiH}_3$  molar ratio ( $= 3\text{S}_{2\text{H}}/2\text{S}_{3\text{H}}$ ). Typically, while the  $\text{SiH}_2/\text{SiH}_3$  molar ratio is 5.0-19.0 in the case of the raw material polysilazane, the polycondensed product gives a reduced  $\text{SiH}_2/\text{SiH}_3$  molar ratio of 2.5-8.4.

Since the polycondensed polysilazane of the present invention is soluble in various organic solvents and since it is able to be converted into silicon nitride or silicon nitride-containing ceramics upon calcination, shaped ceramic bodies such as continuous fibers, films and coatings which have good properties such as mechanical strengths at elevated temperatures, heat resistance, corrosion resistance, oxidation resistance and resistance to thermal shock may be obtained from the reformed polysilazane. The high ceramic yield of the polycondensed polysilazane also permits the use thereof as binders and

impregnating materials. Moreover, the polycondensed polysilazane, which is free of undesirable impurities such as metal catalysts, catalysts causing decomposition of the polysilazane, or the like impurities, is stable and easy to handle, withstands a long term storage, and gives ceramics of a high purity with a high ceramic yield. Additionally, because of the high molecular weight and the increased crosslinkage of the polycondensed polysilazane, the solidifiability, moldability and the spinnability are improved. Further, the polycondensed polysilazane can be easily obtained by the method of the present invention.

The following examples will further illustrate the present invention.

#### Example 1

To a four-necked 1000 ml flask equipped with a gas feed conduit, a mechanical stirrer and a Dewar condenser, an oxygen-free, dry nitrogen gas was fed to replace the inside of the flask therewith. After charging 490 ml of dry pyridine (deaerated), the flask was cooled in an ice bath. Then 51.6 g of dichlorosilane were added into the flask to form a white precipitate of adduct ( $\text{SiH}_2\text{Cl}_2 \cdot 2\text{C}_5\text{H}_5\text{N}$ ). Subsequently, with stirring and cooling in the ice bath, 51.0 g of ammonia which had been passed successively through a sodium hydroxide-containing tube and an active carbon-containing tube was bubbled through the reaction mixture within the flask. After completion of the reaction, the reaction mixture was centrifuged and the supernatant was washed with dry pyridine, followed by filtration in a nitrogen atmosphere to give 850 ml of a filtrate containing perhydropolysilazane. When the solvent was removed from the filtrate (5 ml) by evaporation in vacuo, 0.102 g of resinous solid was obtained. Gel permeation chromatography of the polysilazane revealed that the polysilazane had a number-average molecular weight of 980. The IR spectrum of the polysilazane (solvent: dry o-xylene, concentration: 10.2 g/liter) showed absorptions at  $3350\text{ cm}^{-1}$  (absorbance index  $\epsilon = 0.557\text{ liter}\cdot\text{g}^{-1}\cdot\text{cm}^{-1}$ ) and  $1175\text{ cm}^{-1}$  (based on NH), at  $2170\text{ cm}^{-1}$  ( $\epsilon = 3.14$ , based on SiH) and at  $1020\text{--}820\text{ cm}^{-1}$  (based on SiH and SiNSi). The 60 MHz proton NMR spectrum (solvent:  $\text{CDCl}_3$ , reference: TMS) showed broad signals at 4.8 ( $\text{SiH}_2$ ) and 4.4 ( $\text{SiH}_3$ ) and at 1.5 (NH). This polysilazane will be hereinafter referred to as polysilazane (I).

Polysilazane (I) was dissolved in pyridine to give 100 ml of a pyridine solution containing 10.3% by weight of polysilazane (I). The solution was charged in an autoclave with a 300 ml inside volume, to which 3 ml (29.0 mmol) of dehydrated diethylamine were added. The mixture was then reacted at  $110^\circ\text{C}$  under autogenous pressure for 3 hours with stirring. A large amount of a gas was found to be produced and the pressure within the autoclave was increased by  $2.2\text{ kg/cm}^2$  as a result of the reaction. The gas evolved was found to be hydrogen by gas chromatography. After cooling to room temperature, dry o-xylene (200 ml) was mixed with the reaction mixture. The solvent was then removed by distillation at  $50\text{--}70^\circ\text{C}$  under 3-5 mmHg to give 5.1 g of white powder. This powder (polycondensed product) was found to be soluble in toluene, tetrahydrofuran and chloroform. Gel permeation chromatography revealed that the polycondensed product had a number-average molecular weight of 5240. The IR spectrum of the polycondensed product (solvent: toluene) showed absorptions at  $3350\text{ cm}^{-1}$  and  $1175\text{ cm}^{-1}$  (based on NH), at  $2170\text{ cm}^{-1}$  (based on SiH), at  $1020\text{--}820\text{ cm}^{-1}$  (based on SiH and SiNSi) at 2690, 2930, 2880, 1460 and  $1380\text{ cm}^{-1}$  (based on CH) and at  $1090\text{ cm}^{-1}$  (based on CN). The proton NMR spectrum (solvent:  $\text{CDCl}_3$ , reference: TMS) showed broad signals at 4.8 ( $\text{SiH}_2$ ) at 4.4 ( $\text{SiH}_3$ ), at 2.7 ( $\text{CH}_2$ ), at 1.4 (NH) and at 0.9 ( $\text{CH}_3$ ). The elementary analysis of the polycondensed product gave (% by weight):

Si: 50.1, N: 25.8, O: 3.6, C: 12.4, H: 7.3

#### Example 2

Polysilazane (I) obtained in Example 1 was dissolved in  $\gamma$ -picoline to give 80 ml of a pyridine solution containing 5.76% by weight of polysilazane (I). The solution was charged in an autoclave with a 300 ml inside volume, to which 5 ml of trimethylhydrazine were added. The mixture was then reacted at  $120^\circ\text{C}$  under autogenous pressure for 3 hours with stirring. A large amount of a gas was found to be produced and the pressure within the autoclave was increased by  $2.0\text{ kg/cm}^2$  as a result of the reaction.

After cooling to room temperature, dry o-xylene (200 ml) was mixed with the reaction mixture. The solvent was then removed by distillation at  $50\text{--}70^\circ\text{C}$  under 3-5 mmHg to give 4.9 g of white powder which was soluble in toluene, tetrahydrofuran and chloroform. Gel permeation chromatography revealed that this powder (polycondensed polysilazane) had a number-average molecular weight of 3080.

Example 3

To a four-necked 1000 ml flask equipped with a gas feed conduit, a mechanical stirrer and a Dewar condenser, an oxygen-free, dry nitrogen gas was fed to replace the inside of the flask therewith. After charging 500 ml of dry dichloromethane, the flask was cooled in an ice bath. Then 48.6 g of dichlorosilane were added into the flask. Subsequently, with stirring and cooling in the ice bath, 42.5 g of ammonia which had been passed successively through a sodium hydroxide-containing tube and an active carbon-containing tube was, as a mixture with dry nitrogen gas, bubbled through the reaction mixture within the flask. During the course of the reaction, it was necessary to occasionally beat the gas passage downstream of the flask in order to remove a powdery substance which deposited on and accumulated in the inside surface thereof and to prevent the clogging thereof. After completion of the reaction, the reaction mixture was centrifuged and the supernatant was washed with dry pyridine. After filtration, the solvent was removed from the filtrate by evaporation in vacuo to leave 9.6 g of viscous oil polysilazane. Gel permeation chromatography of the polysilazane revealed that the polysilazane had a number-average molecular weight of 640. This polysilazane will be hereinafter referred to as polysilazane (II).

Polysilazane (II) was dissolved in pyridine to give 100 ml of a pyridine solution containing 7.04% by weight of polysilazane (II). The solution was charged in an autoclave with a 300 ml inside volume, to which 25 ml of N-methylaniline were added. The mixture was then reacted at 100 °C under autogenous pressure for 6 hours with stirring. A large amount of a gas was found to be produced so that the pressure within the autoclave was increased by 1.8 kg/cm<sup>2</sup> as a result of the reaction. After cooling to room temperature, o-xylene was mixed with the reaction mixture. The solvent was then removed by distillation at 50-70 °C under 3-5 mmHg to give 6.8 g of light yellow powder which was soluble in toluene, tetrahydrofuran and chloroform. The gel permeation chromatography revealed that this powder (polycondensed polysilazane) had a number-average molecular weight of 2740.

Example 4

To a four-necked 1000 ml flask equipped with a gas feed conduit, a mechanical stirrer and a Dewar condenser, an oxygen-free, dry nitrogen gas was fed to replace the inside of the flask therewith. After charging 450 ml of deaerated dry tetrahydrofuran, the flask was cooled in a dry ice-methanol bath. Then 46.2 g of dichlorosilane were added into the flask. Subsequently, with stirring and cooling, 44.2 g of anhydrous methylamine was, as a mixture with dry nitrogen gas, bubbled through the reaction mixture within the flask. After completion of the reaction, the reaction mixture was centrifuged and the supernatant was washed with dry tetrahydrofuran, followed by filtration in a nitrogen atmosphere to obtain 820 ml of a filtrate. The solvent was removed from the filtrate by evaporation in vacuo to leave 8.4 g of viscous oil N-methylsilazane. The gel permeation chromatography revealed that the N-methylsilazane had a number-average molecular weight of 1100. This N-methylsilazane will be hereinafter referred to as polysilazane (III).

Polysilazane (III) was dissolved in pyridine to give 60 ml of a pyridine solution containing 12.4% by weight of polysilazane (III). The solution was charged in an autoclave with a 300 ml inside volume, to which 20 g of 1,2-dimethyl-1-phenylhydrazine were added. The mixture was then reacted at 150 °C under autogenous pressure for 5 hours with stirring. A gas was found to be produced so that the inside pressure was increased by 0.3 kg/cm<sup>2</sup> as a result of the reaction. After cooling to room temperature, o-xylene was mixed with the reaction mixture. The solvent was then removed by distillation at 50-70 °C under 3-5 mmHg to give 6.3 g of light yellow, wax-like solid having a melting point of about 200 °C. Gel permeation chromatography revealed that this product (polycondensed polysilazane) had a number average molecular weight of 1480.

Example 5

To a four-necked 1000 ml flask equipped with a gas feed conduit, a mechanical stirrer and a Dewar condenser, an oxygen-free, dry nitrogen gas was fed to replace the inside of the flask therewith. After charging 300 ml of dry dichloromethane and 24.3 g (0.211 mol) of methylchlorosilane, the flask was cooled in an ice bath. Then, with stirring and cooling in the ice bath, 18.1 g of ammonia which had been passed successively through a sodium hydroxide-containing tube and an active carbon-containing tube for purification was bubbled through the reaction mixture within the flask. After completion of the reaction, the reaction mixture was centrifuged and the supernatant was washed with dry dichloromethane, followed by filtration in a nitrogen atmosphere. The solvent was removed from the filtrate by evaporation in vacuo to leave 8.81 g of polymethylsilazane in the form of a colorless, transparent liquid. Gel permeation chromatog-

raphy of the polysilazane revealed that the polysilazane had a number-average molecular weight of 380. This polysilazane will be hereinafter referred to as polysilazane (IV).

Polysilazane (IV) was dissolved in pyridine to give 100 ml of a solution containing 8.16% by weight of polysilazane (III). The solution was charged in an autoclave with a 300 ml inside volume, to which 15 ml of dipropylamine were added. The mixture was then reacted at 120 °C under autogenous pressure for 6 hours with stirring. A gas was found to be produced so that the inside pressure was increased by 0.4 kg cm<sup>-2</sup> as a result of the reaction. After cooling to room temperature, the solvent was removed by distillation at 50-70 °C under 3-5 mmHg to give 7.8 g of white wax-like solid. Gel permeation chromatography revealed that this product (polycondensed polysilazane) had a number-average molecular weight of 1720.

#### Example 6

To a four-necked 1000 ml flask equipped with a gas feed conduit, a mechanical stirrer and a Dewar condenser, an oxygen-free, dry nitrogen gas was fed to replace the inside of the flask therewith. After charging 500 ml of dry benzene, the flask was cooled in a water bath. Then 40.6 g of dichlorosilane were added into the flask. Subsequently, with stirring and cooling in the water bath, 42.0 g of ammonia which had been passed successively through a sodium hydroxide-containing tube and an active carbon-containing tube for purification was, as a mixture with dry nitrogen gas, bubbled through the reaction mixture within the flask. During the course of the reaction, it was necessary to occasionally beat the gas passage downstream of the flask in order to remove a powdery substance which deposited on and accumulated in the inside surface thereof and to prevent the clogging thereof. After completion of the reaction, the reaction mixture was centrifuged and the supernatant was washed with dry o-xylene. After filtration, the solvent was removed from the filtrate by evaporation in vacuo to leave 5.2 g of viscous oily perhydropolysilazane. Gel permeation chromatography of the polysilazane revealed that the polysilazane had a number-average molecular weight of 320. This polysilazane is referred to as polysilazane (V) hereinafter.

Polysilazane (V) was dissolved in pyridine to give 100 ml of a pyridine solution containing 5.59% by weight of polysilazane (V). The solution was charged in an autoclave with a 300 ml inside volume, to which 8 ml of trimethylhydrazine were added. The mixture was then reacted at 120 °C under autogenous pressure for 3 hours with stirring. A large amount of a gas was found to be produced so that the pressure within the autoclave was increased by 0.8 kg/cm<sup>2</sup> as a result of the reaction. After cooling to room temperature, the solvent was removed by distillation at 50-70 °C under 3-5 mmHg to give 5.2 g of white powder. The gel permeation chromatography revealed that this powder (polycondensed polysilazane) had a number-average molecular weight of 2250.

#### Example 7

To a four-necked 1000 ml flask equipped with a gas feed conduit, a dropping funnel, a mechanical stirrer and a Dewar condenser, an oxygen-free, dry nitrogen gas was fed to replace the inside of the flask therewith. Then, 400 ml of dry benzene and 64.5 g of allyldichlorosilane which had been obtained by the conventional method (J. Am. Chem. Soc., Vol. 67, 1813 (1945)) were added into the flask. While stirring the mixture in the flask, a solution of 42.5 g of triethylaminosilane in 50 ml dry benzene was added dropwise from the dropping funnel. The triethylaminosilane used was obtained by the known method (J. Am. Chem. Soc., Vol. 70, 435 (1948)). Thereafter, the reaction mixture was refluxed in an oil bath with stirring to complete the reaction. The reaction mixture was centrifuged and the supernatant was washed with dry benzene. After filtration in a nitrogen atmosphere, the solvent was removed from the filtrate by evaporation in vacuo to leave 19.2 g of N-(triethylsilyl)allylsilazane in the form of liquid. Gel permeation chromatography of the polysilazane revealed that the polysilazane had a number-average molecular weight of 360. This polysilazane is referred to as polysilazane (VI) hereinafter.

Polysilazane (VI) was dissolved in pyridine to give 60 ml of a pyridine solution containing 6.47% by weight of polysilazane (VI). The solution was charged in an autoclave with a 300 ml inside volume, to which 15 ml of diethylamine were added. The mixture was then reacted at 160 °C under autogenous pressure for 6 hours with stirring. A gas was found to be produced so that the pressure within the autoclave was increased by 0.3 kg/cm<sup>2</sup> as a result of the reaction. After cooling to room temperature, the solvent was removed by distillation at 50-70 °C under 3-5 mmHg to give 3.2 g of light yellow, wax-like solid having a melting point of about 180 °C. The gel permeation chromatography revealed that this product (polycondensed polysilazane) had a number-average molecular weight of 1350.



### Example 8

To 110 g of trichlorosilane, 62.0 g of Grignard reagent obtained from cyclohexyl bromide were gradually added. The distillation in vacuo of the mixture gave 16.4 g of cyclohexyldichlorosilane. To a four-necked 1000 ml flask equipped with a gas feed conduit, a dropping funnel, a mechanical stirrer and a Dewar condenser, an oxygen-free, dry nitrogen gas was fed to replace the inside of the flask therewith. Then, 420 ml of dry benzene and 12.5 g of allyldichlorosilane obtained above were added into the flask. While stirring the mixture in the flask, a solution of 15.6 g of 1,1-dimethylhydrazine in 40 ml of dry benzene was added dropwise from the dropping funnel. Thereafter, the reaction mixture was stirred at room temperature to complete the reaction. The reaction mixture was centrifuged and the supernatant was washed with dry benzene. After filtration in a nitrogen atmosphere, the solvent was removed from the filtrate (730 ml) by evaporation in vacuo to leave 3.2 g of N-(dimethylamino)-cyclohexylsilazane in the form of an oil. Gel permeation chromatography of the polysilazane revealed that the polysilazane had a number-average molecular weight of 390. This polysilazane is referred to as polysilazane (VII) hereinafter.

Polysilazane (VII) was dissolved in pyridine to give 80 ml of a pyridine solution containing 3.2% by weight of polysilazane (VII). The solution was charged in an autoclave with a 300 ml inside volume, to which 10 ml of dibutylamine were added. The mixture was then reacted at 120 °C under autogenous pressure for 8 hours with stirring. A gas was found to be produced so that the pressure within the autoclave was increased by 0.2 kg/cm<sup>2</sup> as a result of the reaction. After cooling to room temperature, the solvent was removed by distillation at 50-70 °C under 3-5 mmHg to give 2.0 g of light yellow, wax-like solid having a melting point of about 150 °C. The gel permeation chromatography revealed that this product (polycondensed polysilazane) had a number-average molecular weight of 1110.

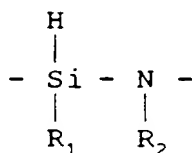
### Example 9

To a four-necked 1000 ml flask equipped with a gas feed conduit, a mechanical stirrer and a Dewar condenser, an oxygen-free, dry nitrogen gas was fed to replace the inside of the flask therewith. After charging 500 ml of dry toluene, the flask was cooled in an ice bath. Then 52.1 g of phenyldichlorosilane were added into the flask. Subsequently, with stirring and cooling in the ice bath, 30.0 g of ammonia which had been passed successively through a sodium hydroxide-containing tube and an active carbon-containing tube was bubbled through the reaction mixture within the flask. After completion of the reaction, the reaction mixture was centrifuged and the supernatant was washed with dry pyridine, followed by filtration in a nitrogen atmosphere. The solvent was removed from the filtrate by evaporation in vacuo to leave 6.8 g of oily phenylpolysilazane. Gel permeation chromatography revealed that the polysilazane had a number-average molecular weight of 380. This polysilazane will be hereinafter referred to as polysilazane (VIII).

Polysilazane (VIII) was dissolved in pyridine to give 100 ml of a pyridine solution containing 8.14% by weight of polysilazane (VIII). The solution was charged in an autoclave with a 300 ml inside volume, to which 30 ml of dipropylamine were added. The mixture was then reacted at 150 °C under autogenous pressure for 6 hours with stirring. A gas was found to be produced and the pressure within the autoclave was increased by 0.4 kg/cm<sup>2</sup> as a result of the reaction. After cooling to room temperature, the solvent was removed by distillation at 50-70 °C under 3-5 mmHg to give 7.3 g of white wax-like solid. This product (polycondensed polysilazane) was found to have a number-average molecular weight of 1310 upon analysis by gel permeation chromatography.

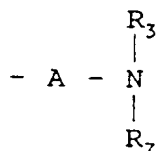
### Claims

1. A polycondensed polysilazane having a number average molecular weight of 200-500,000, and containing (a) recurring units of the formula:



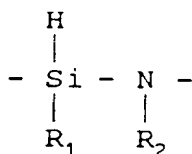
wherein R<sub>1</sub> and R<sub>2</sub>, independently from each other, stand for hydrogen, alkyl, alkenyl, cycloalkyl,

alkylamino, aryl or alkylsilyl, and (b) one or more groups replacing hydrogen on and pendant from the silicon atoms of said repeating units, of the formula:



wherein A stands for a direct bond or  $-\text{N}(\text{R}_4)-$  where  $\text{R}_4$  stands for hydrogen, alkyl, cycloalkyl, aryl, aralkyl or a heterocyclic group, and  $\text{R}_3$  and  $\text{R}_7$ , independently from each other, stand for alkyl, alkenyl, cycloalkyl, aryl, aralkyl or a heterocyclic group.

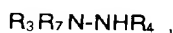
2. A polysilazane according to Claim 1, wherein the number average molecular weight is 1,000-50,000.
3. A method of forming a novel polycondensed polysilazane as defined in Claim 1, comprising reacting a starting polysilazane having a number average molecular weight of 100-50,000 and containing recurring units of the formula:



wherein  $\text{R}_1$  and  $\text{R}_2$  are as defined in Claim 1, in an organic base-containing solvent, with a secondary amine of the formula:



or with a 1,1,2-trisubstituted hydrazone of the formula:



wherein  $\text{R}_3$ ,  $\text{R}_4$  and  $\text{R}_7$  are as defined in Claim 1.

4. A method according to Claim 3, wherein the organic base-containing solvent consists of a liquid organic base.
5. A method according to Claim 4, wherein the liquid organic base is a member selected from the group consisting of trimethylamine, dimethylamine, dimethylethylamine, diethylmethylamine, triethylamine, pyridine and a substituted pyridine, dimethylaniline and a substituted dimethylaniline, pyrazine and a substituted pyrazine, pyrimidine and a substituted pyrimidine, pyridazine and a substituted pyridazine, pyrrole, 3-pyrroline, pyrazole, 2-pyrazoline and mixtures thereof.
6. A method according to Claim 3, wherein the organic base-containing solvent is a non-basic organic solvent having dissolved therein an organic base.
7. A method according to Claim 6, wherein said organic base is a member selected from the group consisting of tertiary amines, secondary amines, phosphines and mixtures thereof and said non-basic organic solvent is a member selected from the group consisting of methylene chloride, chloroform, carbon tetrachloride, bromoform, ethylene chloride, ethylidene chloride, trichloroethane, tetrachloroethane, ethyl ether, isopropyl ether, ethyl butyl ether, butyl ether, 1,2-dioxoethane, dioxane, dimethyl dioxane, tetrahydrofuran, tetrahydropyran, pentane, hexane, isohexane, methylpentane, heptane, isohexane, octane, isooctane, cyclopentane, methylcyclopentane, cyclohexane, methylcyclohexane, benzene, toluene, xylene, ethylbenzene and mixtures thereof.

8. A method according to any one of Claims 3-7, wherein said polysilazane solution has a concentration of 0.1-50% by weight.
- 5 9. A method according to any one of Claims 3-8, wherein said compound of the formula (I) is used in an amount of 0.001-2000 mols per mole of said polysilazane when the compound of the formula (I) is ammonia or hydrazine and in an amount of 0.01-5000 mols per mole of said polysilazane when said compound of the formula (I) is an amine or a substituted hydrazine.
- 10 10. A method according to any one of Claims 3-9, wherein said starting polysilazane has a number-average molecular weight of 100-50,000 and said reaction is performed so that the resulting polysilazane has a number-average molecular weight of 200-500,000.
- 15 11. A method according to any one of Claims 3-10, wherein said reaction is performed so as to increase the N/Si molar ratio of said polysilazane.

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(54) **Polycondensed polysilazane and method of producing same.**

(57) A novel, polycondensed polysilazane is obtained by reacting a polysilazane with a compound selected from secondary amines and tri-substituted hydrazines so as to link the compound to the polysilazane.

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Office

# EUROPEAN SEARCH REPORT

Application Number

EP 93 20 1106

## DOCUMENTS CONSIDERED TO BE RELEVANT

| DOCUMENTS CONSIDERED TO BE RELEVANT   |   |   |   |
|---|---|---|---|
| Category  | Citation of document with indication, where appropriate, of relevant passages   | Relevant to claim   | CLASSIFICATION OF THE APPLICATION (Int. Cl.4) |
| A   | US-A-4 659 850 (M. ARAI ET AL.)<br>* claims *   | 1-11  | C08G77/62                                     |
| A   | WO-A-8 606 377 (SRI INTERNATIONAL)<br>* page 1, lines 20-24;<br>page 5, line 28 - page 6, line 14;<br>examples 1, 5, 11, 13, 17, 18;<br>claims 18-20, 24, 26-29 * | 1-11  |   |
| D,A   | US-A-4 482 669 (D. SEYFERTH ET AL.)<br>* claims 1-3 *   | 1,2   |   |
| A   | US-A-4 612 383 (R. M. LAINE ET AL.)<br>* abstract; examples 1, 5, 6;<br>claim 1 *   | 1-3   |   |
|   |   |   | TECHNICAL FIELDS SEARCHED (Int. Cl.4)         |
|   |   |   | C08G  |
| The present search report has been drawn up for all claims  |   |   |   |
| Place of search   |   | Date of completion of the search  | Examiner                                      |
| BERLIN  |   | 02 JULY 1993  | HOEPFNER W.                                   |
| CATEGORY OF CITED DOCUMENTS   |   |   |   |
| X : particularly relevant if taken alone<br>Y : particularly relevant if combined with another document of the same category<br>A : technological background<br>O : non-written disclosure<br>P : intermediate document |   | T : theory or principle underlying the invention<br>E : earlier patent document, but published on, or after the filing date<br>D : document cited in the application<br>L : document cited for other reasons<br>-----<br>& : member of the same patent family, corresponding document |   |